COMPLEXING IN BIOLOGICAL

SYSTEMS

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<u>Coordination compounds</u> are the higher-order compounds which are stable in aqueous solutions or dissociate insignificantly.



Central atom (ion) or complexing agent takes central place in coordination compounds and is usually a positively charged ion



Ligands are ions of opposite charge or neutral molecules that are located (coordinated) around the complexing agent.

CN⁻

K₄[Fe(CN)₆]

Inner sphere (coordination entity) is formed by complexing agent and ligands. $[Fe(CN)_6]^{4-}$

Outer sphere is formed by lons which are not included in the inner sphere.

 K^+

Depending on the electric charge of the inner sphere

- Coordination compounds containing <u>complex cations</u> -[Zn(NH3)4]Cl2
- 2. Coordination compounds containing <u>complex anions</u> $K_3[AI(OH)_6]$
- 3. <u>Neutral complexes</u> [Pt(NH₃)₂Cl₂]

- Depending on the nature of ligands:
- 1. <u>Acid complexes</u> $[Fe(CN)_6]^{4-}$
- 2. Aqua complexes $[Cr(H_2O)_6]^{3+}$
- 3. <u>Hydroxide complexes</u> $[Zn(OH)_4]^{2-}$
- 4. <u>Ammonia complexes</u> $[Cu(NH_3)_4]^{2+}$

- Depending on the chemical compounds class:
- 1. Acids $H[AuCl_4]$
- 2. Bases [Ag(NH₃)₂]OH
- 3. Salts $K_2[Hgl_4]$

- Depending on the quantity of central atoms:
- 1. Mononuclear $[Cr(NH_3)_3(H_2O)_3]Cl_3$
- 2. Polynuclear [Pt4(OH)4](ClO4)4

When didentate or polydentate ligand uses its two or more donor atoms to bind a single metal ion, it is said to be *a chelate ligand*

The coordination compounds in which a ligand is bound with central atom by both donor-acceptor bond and ionic bond are called

chelates

CHELATES APPLICATION IN MEDICINE

Disease	Excess metal ion	Used chelating agent
Hemochromatosis, hemosiderosis, iron intoxication	Fe	Deferoxamine, Penicillamine
Cataracts, atherosclerosis	Ca	Na ₂ EDTA , Penicillamine
Wilson's disease	Cu	Mixture of Penicillamine and Calcium Tetacine
Disease "itai-itai-bio"	Cd	Cryptand, Calcium Tetacine
Minimata disease	Hg	Penicillamine, Calcium Tetacine
Plutonium poisoning	Pu	Pentacine
Lead poisoning	Pb	Calcium Tetacine, Na2EDTA
Berylliosis, beryllium rickets	Be	Aluminon