

Лекция 5

Электронная система поиска Скопус. Особенности работы

С.М. Пестов

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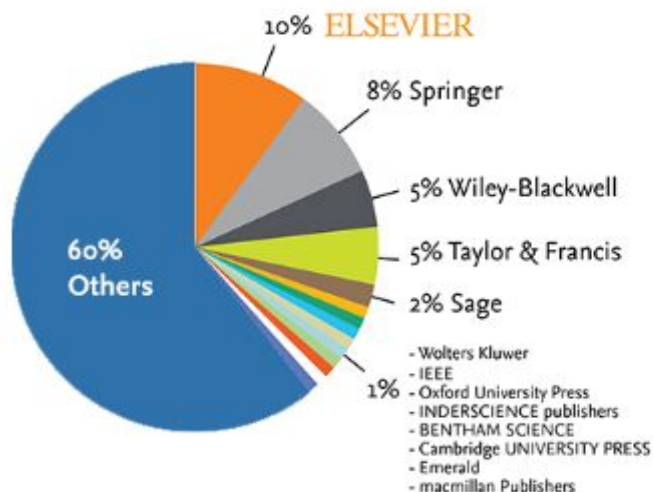
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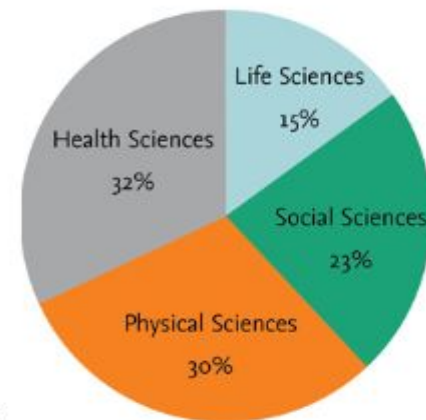
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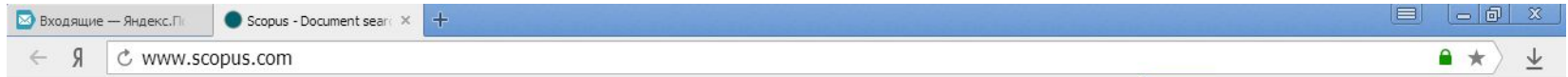


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adsorption of copper along with **gold** causes difficulty in separating **gold** and copper at the **gold** elution stage. Our previous study has demonstrated that nickel catalyzed ammonium thiosulfate solution for **gold** extraction has the advantage of reducing thiosulfate consumption. In this study, the results also demonstrated the advantage of **gold recovery** from the nickel catalyzed ammonium thiosulfate solution by strongly basic anion **exchange resin**. The optimal **gold** loading conditions on a 1 g/dm³ strongly base anion **exchange resin** (wet base value) are investigated in several **ion** concentrations and 95 kg-Au/t-**resin** has been obtained. The alternative **gold** eluant was investigated as the **gold** loaded **resin** cannot be eluted by conventional hydrochloric acid. Results showed that the elution efficiency was in the order of OH⁻ < Cl⁻ < NO₃⁻ < Br⁻ < I⁻ < ClO₄⁻. The maximum **gold recovery** by using 2.5 mol/dm³ ClO₄⁻ was around 98% with the stripped **resin** assayed as 0.2kg/tAu. The feasibility of **resin** recycling has demonstrated that there was no deterioration in **gold** adsorption and desorption for four cycles.

Leaching and recovery of gold using ammoniacal thiosulfate leach liquors (a review) 65 Grosse, A.C., Dicinovski, G.W., Shaw, M.J., Haddad, P.R. 2003 Hydrometallurgy 69 (1-3), pp. 1-21 92 Cited by

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A review is presented summarising the leaching of **gold** with ammoniacal thiosulfate solutions, and evaluating the current use and development of **ion exchange resins** for the **recovery of gold** and silver from such leach liquors. Comparisons are also made with other **recovery** processes, including carbon adsorption, solvent extraction, electrowinning and precipitation. Thiosulfate leaching chemistry is compared with cyanide leaching, and the problems associated with obtaining a high yield of recovered **gold** using the former process are discussed. The present limitations of using **Resin-in-Pulp (RIP)** and **Resin-in-Leach (RIL)** systems with thiosulfate liquors are indicated and possible solutions discussed. © 2002 Elsevier Science B.V. All rights reserved.

Recent advances in the development of an alternative to the cyanidation process: Thiosulfate leaching and resin in pulp 66 Fleming, C.A., McMullen, J., Thomas, K.G., Wells, J.A. 2003 Minerals and Metallurgical Processing 39

A process based on thiosulfate leaching followed by **resin-in-pulp gold** extraction was developed to treat the carbonaceous, preg-robbing ores of Barrick's Goldstrike orebody in the Carlin Trend of Nevada. These ores have proven to be amenable to thiosulfate leaching under mild conditions. **Gold** leaches rapidly as the **gold** thiosulfate complex, which, because of its low affinity for graphitic carbon, does not suffer the preg-robbing phenomenon that is a feature of these ores in cyanide leach circuits. The mild leaching conditions are also compatible with a **gold-recovery** process involving direct **recovery** from the leach pulp by adsorption on strong-base anion-**exchange resins**. Finally, a novel elution/regeneration process was developed to elute the **gold** off the **resin**, recover the **gold** from the eluate and restore the **resin** for recycling to the adsorption circuit.

Determination of trace gold by on-line enrichment flow injection flame atomic absorption spectrometry with N1923 levetrel resin 67 Ye, M.D., Xue, X.Y. 2003 Guang pu xue yu guang pu fen xi = Guang pu 1

A new method for the determination of micro amount of **gold** with N1923 levetrel **resin** by flow injection on-line separation and flame atomic absorption spectrometry is described. Au (III) absorbed on the **resin** can be eluted quantitatively using sulphuric-urea solution. The absorption is carried out in 1.0 mol.L⁻¹ HCl medium and the enhancement factor of 32 is achieved for a loading period of 90 s. The detection limit is 0.001 microgram.mL⁻¹. The flow rate of sample injection, the time of extraction, the flow rate of enrichment, the concentration and acidity of eluting and the effect of coexistence element are studied by flame atomic absorption spectrometry. The **recoveries** of Au are 98.3%-101%. The developed method has been applied to the determination of trace **gold** in water samples with satisfactory results.

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Mineral Processing and Extractive Metallurgy Review

Volume 37, Issue 2, 3 March 2016, Pages 73-119

Heap leaching technology - Current State, innovations, and future directions: A review (Article)

Ghorbani, Y.^{ab}, Franzidis, J.-P.^c, Petersen, J.^c

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^c Minerals to Metals Signature Theme, Department of Chemical Engineering, University of Cape Town, Private Bag X6, Rondebosch, South Africa

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Abstract

Heap leaching is a well-established extractive metallurgical technology enabling the economical processing of various kinds of low-grade ores, which could not otherwise be exploited. However, despite much progress since it was first applied in recent times, the process remains limited by low recoveries and long extraction times. It is becoming increasingly clear that the choice of heap leaching as a suitable technology to process a particular mineral resource, which is both environmentally sound and economically viable, very much depends on having a comprehensive understanding of the underlying fundamental mechanisms of the processes and how they interact with the particular mineralogy of the ore body under consideration. This paper provides an introduction to the theoretical background of various heap leach processes, offers a scientific and patent literature overview on technology developments in commercial heap leaching operations around the world, identifies factors that drive the selection of heap leaching as a processing technology, describes challenges to exploiting these innovations, and concludes with a discussion on the future of heap leaching. © 2016 Taylor & Francis.

Author keywords

Agglomeration; copper; gold; heap leaching; hydrometallurgy; mineralogy

ISSN: 08827508 CODEIN: MPERE Source Type: Journal Original language: English

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(2011) Minerals Engineering

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Fagan, M.A., Sederman, A.J., Harrison, S.T.L.
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Large particle effects in chemical/biochemical heap leach processes - A review
Ghorbani, Y., Becker, M., Mainza, A.
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<input type="checkbox"/> 1	Assessment of commercial acidic ion-exchange resin for ethyl esters synthesis from <i>Acrocomia aculeata</i> (Macaúba) crude oil	Pasa, T.L.B., Souza, G.K., Diório, A., Arroyo, P.A., Pereira, N.C.	2020	Renewable Energy 146, с. 469-476	0
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<input type="checkbox"/> 2	Ciprofloxacin desorption from gel type ion exchange resin: Desorption modeling in batch system and fixed bed column	Staudt, J., Scheufele, F.B., Ribeiro, C., (...), Canevesi, R., Borba, C.E.	2020	Separation and Purification Technology 230,115857	0
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<input type="checkbox"/> 3	A combined treatment method of novel Mass Bio System and ion exchange for the removal of ammonia nitrogen from micro-polluted water bodies	Tabassum, S.	2019	Chemical Engineering Journal 122217	0
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<input type="checkbox"/> 2	Extraction of extracellular polymers from activated sludge using a cation exchange resin	Frølund, B., Palmgren, R., Keiding, K., Nielsen, P.H.	1996	Water Research 30(8), с. 1749-1758	1492
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<input type="checkbox"/> 5	Novel Ion Exchange Chromatographic Method Using Conductimetric Detection	Small, H., Stevens, T.S., Bauman, W.C.	1975	Analytical Chemistry 47(11), с. 1801-1809	1122
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<input type="checkbox"/> 6	Phase modifiers promote efficient production of hydroxymethylfurfural from fructose	Román-Leshkov, Y., Chheda, J.N., Dumesic, J.A.	2006	Science 312(5782), с. 1933-1937	1092
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2	Biosorbents for heavy metals removal and their future	Wang, J., Chen, C.	2009	Biotechnology Advances 27(2), с. 195-226	1374
3	Selective removal of the heavy metal ions from waters and industrial wastewaters by ion-exchange method	Dabrowski, A., Hubicki, Z., Półkościelny, P., Robens, E.	2004	Chemosphere 56(2), с. 91-106	934
4	Review of fluoride removal from drinking water	Mohapatra, M., Anand, S., Mishra, B.K., Giles, D.E., Singh, P.	2009	Journal of Environmental Management 91(1), с. 67-77	438
5	Recovery of gold from secondary sources-A review	Syed, S.	2012	Hydrometallurgy 115-116, с. 30-51	251
6	Taste masking technologies in oral pharmaceuticals: Recent developments and approaches	Sohi, H., Sultana, Y., Khar, R.K.	2004	Drug Development and Industrial Pharmacy 30(5), с. 429-448	244

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Industrial and Engineering Chemistry Research
 Volume 48, Issue 1, 7 January 2009, Pages 388-398

Ion-Exchange resins: A retrospective from industrial and engineering chemistry research (Review)

Alexandratos, S.D. ✉

Department of Chemistry, Hunter College, City University of New York, 695 Park Avenue, New York, NY 10065, United States

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Ion-exchange resins comprise one of the most important scientific developments of the 20th century. Their applicability to water softening, environmental remediation, wastewater treatment, hydrometallurgy, chromatography, biomolecular separations, and catalysis was recognized in numerous publications. The principle of covalently bonding ligands to cross-linked polymer networks became the basis for the area of polymer-supported reagents. The journal Industrial & Engineering Chemistry Research and its predecessors have published some of the most important papers in this field. In celebration of its 100th anniversary, this review provides a retrospective of ion-exchange resins through publications appearing in this journal. © 2009 American Chemical Society.

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[Improving chloride ion penetration resistance of cement mortar by strong base anion exchange resin](#)
Zhao, P., Zhou, L., Bai, M. (2019) *Construction and Building Materials*

[Application of poly\(vinylphenyltrimethylammonium tribromide\) resin as an efficient polymeric brominating agent in the \$\alpha\$ -bromination and \$\alpha\$ -bromoacetalization of acetophenones](#)
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Original papers: The synthesis, constitution, and uses of bakelite
(1909) *Industrial and Engineering Chemistry*, 1 (3), pp. 149-161. Цитировано 82 раз.
doi: 10.1021/ie50003a004
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Soluble, fusible, resinous condensation products of phenols and formaldehyde.1
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Piffen, P.G., Baldassari, L., Gandolfi, O.
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Adsorptive separations of alkylphenols using ion-exchange resins
Anasthas, H.M., Gaikar, V.G.
(1999) *Reactive and Functional Polymers*

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<input type="checkbox"/> 8	On the controllable soft-templating approach to mesoporous silicates	Wan, Y., Zhao, D.	2007	Chemical Reviews 107(7), с. 2821-2860	1704
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Информация по авторам

Просмотреть сведения об авторе Kumar, S #395
Просмотреть анализ результатов по автору
Организация: Raman Research Institute, SCM Group, Bengaluru, India

13 документов, опубликованных автором Kumar, S., соответствуют вашему запросу (Показаны 13 первых результатов)

Title	Authors	Year	Source
Liquid-crystal nanoscience: An emerging avenue of soft self-assembly	Bisoyi, H.K., Kumar, S.	2011	Chemical Society Reviews
Discotic nematic liquid crystals: Science and technology	Bisoyi, H.K., Kumar, S.	2010	Chemical Society Reviews
Triphenylene-based discotic liquid crystals: recent advances	Pal, S.K., Setia, S., Avinash, B.S., Kumar, S.	2013	Liquid Crystals
Discotic liquid crystal-nanoparticle hybrid systems	Kumar, S.	2014	NPG Asia Materials
Liquid crystals in photovoltaics: A new generation of organic photovoltaics	Kumar, M., Kumar, S.	2017	Polymer Journal
A brief review of carbazole-based photorefractive liquid crystalline materials	Manickam, M., Iqbal, P., Belloni, M., Kumar, S., Preece, J.A.	2012	Israel Journal of Chemistry
Functional discotic liquid crystals	Kumar, S.	2012	Israel Journal of Chemistry
The chemistry of bent-core molecules forming nematic liquid crystals	Kumar, S., Gowda, A.N.	2015	Liquid Crystals Reviews
Discotic Liquid Crystals with Graphene: Supramolecular Self-assembly to Applications	Kumar, M., Gowda, A., Kumar, S.	2017	Particle and Particle Systems Characterization

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<http://orcid.org/0000-0002-4550-4814>

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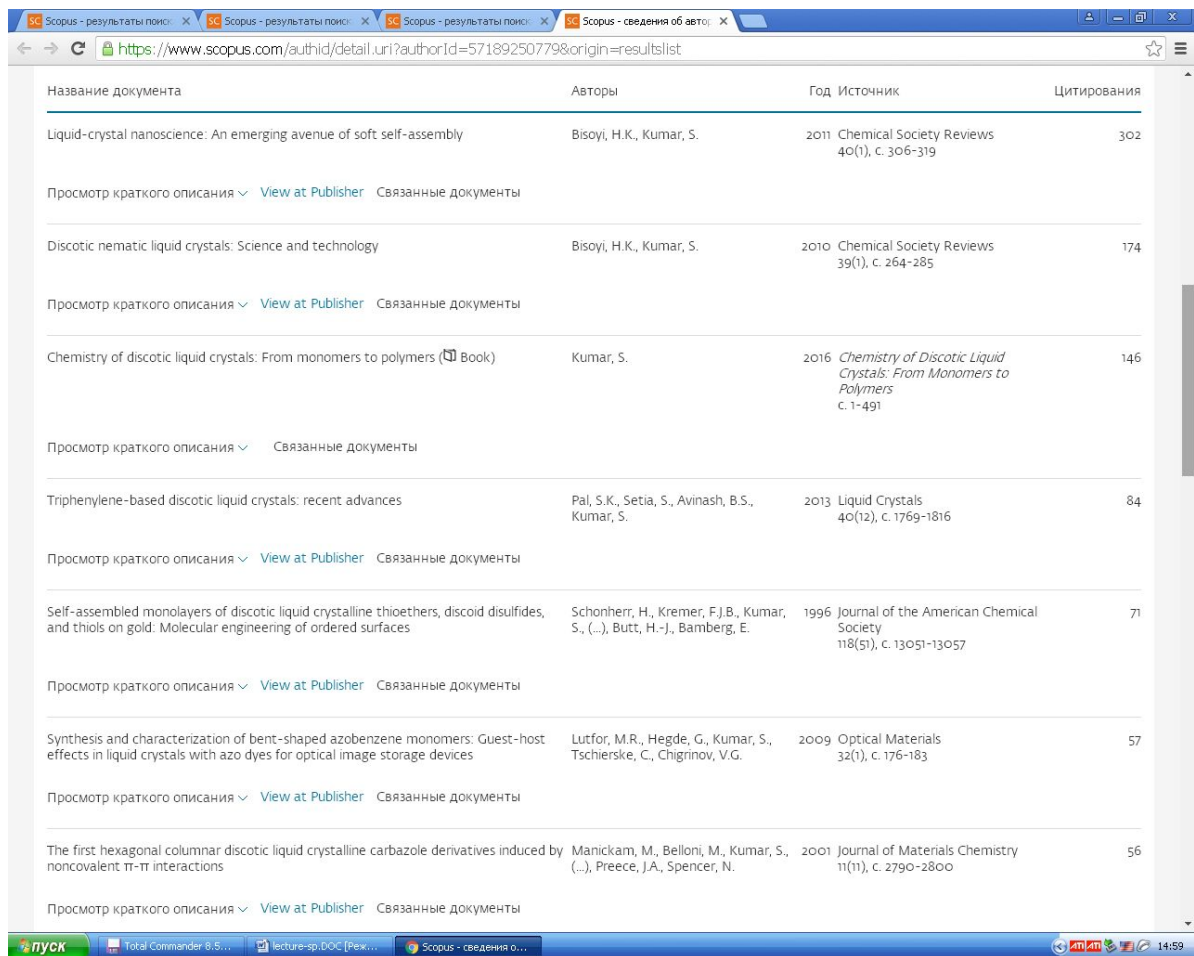
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Название документа	Авторы	Год	Источник	Цитирования
Liquid-crystal nanoscience: An emerging avenue of soft self-assembly	Bisoyi, H.K., Kumar, S.	2011	Chemical Society Reviews 40(1), с. 306-319	302
Discotic nematic liquid crystals: Science and technology	Bisoyi, H.K., Kumar, S.	2010	Chemical Society Reviews 39(1), с. 264-285	174
Chemistry of discotic liquid crystals: From monomers to polymers (Book)	Kumar, S.	2016	Chemistry of Discotic Liquid Crystals: From Monomers to Polymers с. 1-491	146
Triphenylene-based discotic liquid crystals: recent advances	Pal, S.K., Setia, S., Avinash, B.S., Kumar, S.	2013	Liquid Crystals 40(12), с. 1769-1816	84
Self-assembled monolayers of discotic liquid crystalline thioethers, discoid disulfides, and thiols on gold: Molecular engineering of ordered surfaces	Schonherr, H., Kremer, F.J.B., Kumar, S., (...), Butt, H.-J., Bamberg, E.	1996	Journal of the American Chemical Society 118(51), с. 13051-13057	71
Synthesis and characterization of bent-shaped azobenzene monomers: Guest-host effects in liquid crystals with azo dyes for optical image storage devices	Lutfor, M.R., Hegde, G., Kumar, S., Tschierske, C., Chigrinov, V.G.	2009	Optical Materials 32(1), с. 176-183	57
The first hexagonal columnar discotic liquid crystalline carbazole derivatives induced by noncovalent π - π interactions	Manickam, M., Belloni, M., Kumar, S., (...), Preece, J.A., Spencer, N.	2001	Journal of Materials Chemistry 11(11), с. 2790-2800	56

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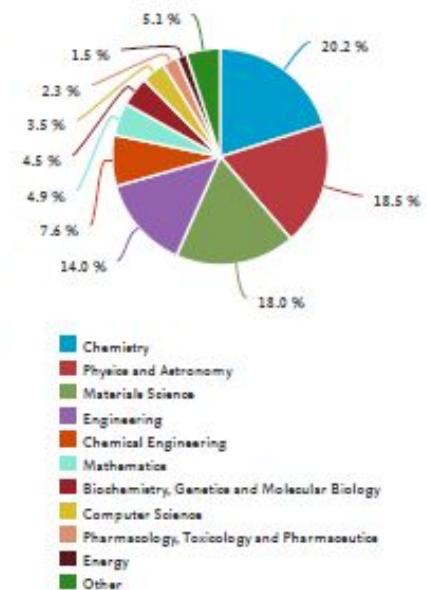
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2013	140
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