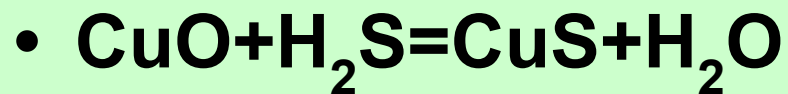
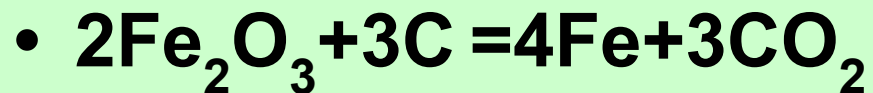
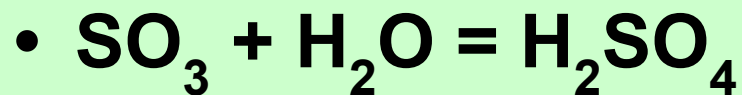
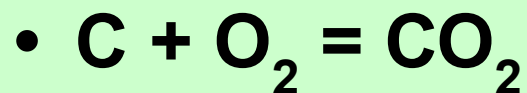


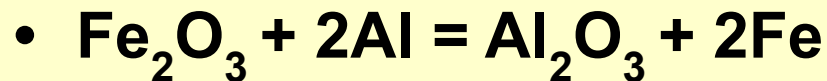
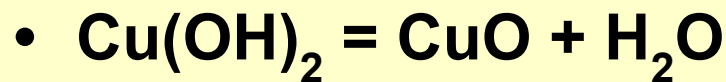
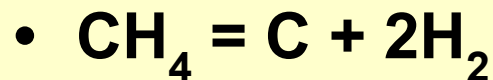
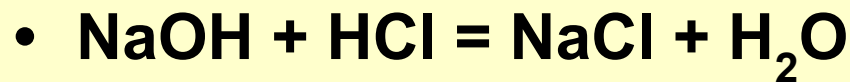
# Составление уравнений ОВР

# Какие из реакций являются окислительно-восстановительными? Укажите окислитель и восстановитель

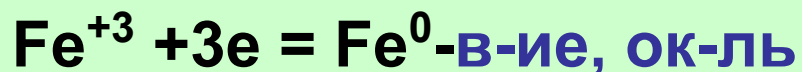
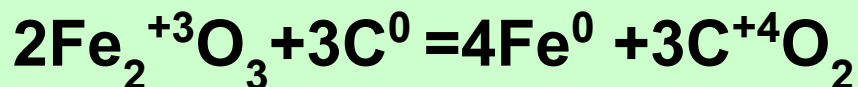
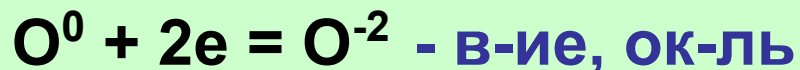
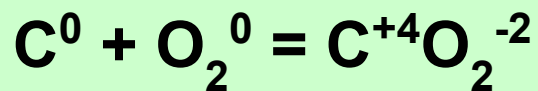
## • 1 Вариант



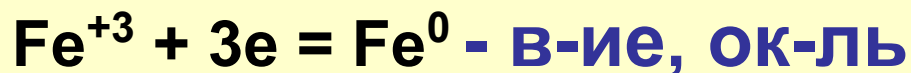
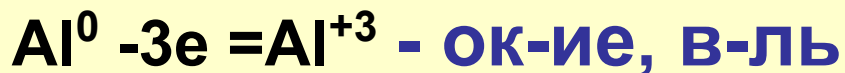
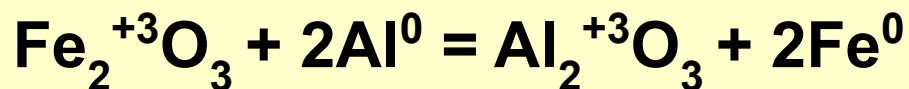
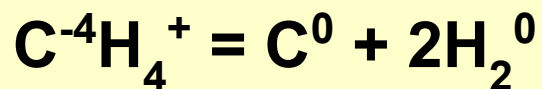
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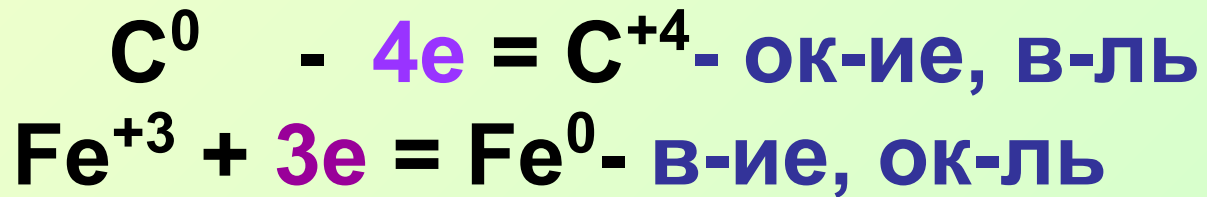
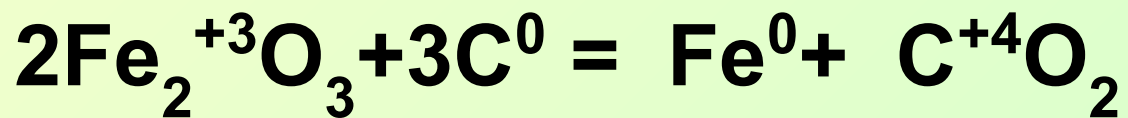


- 1 Вариант



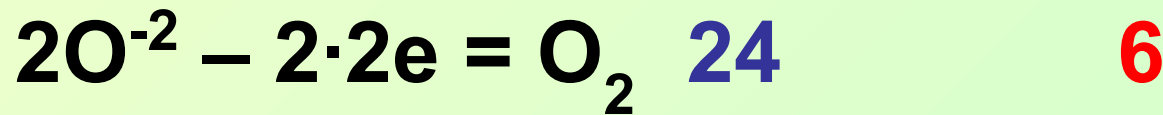
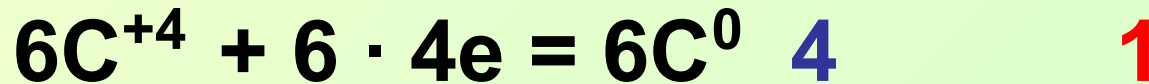
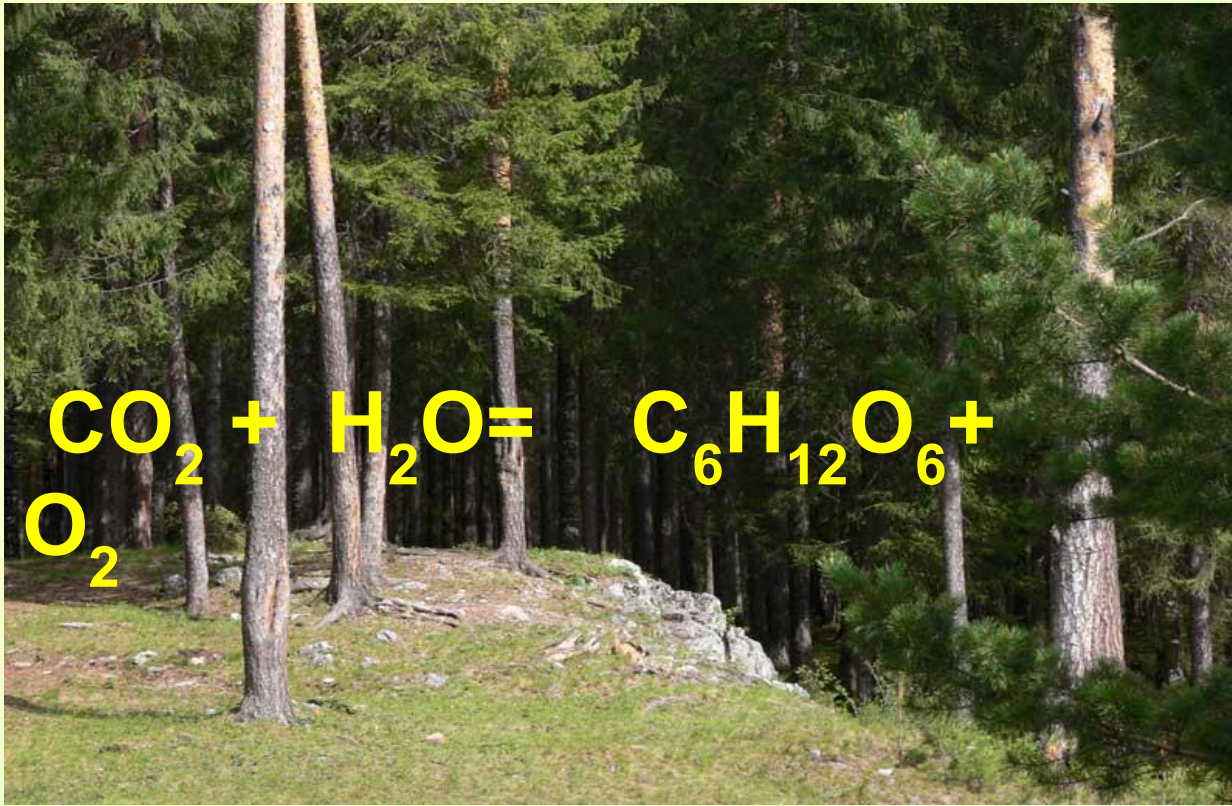
- 2 Вариант



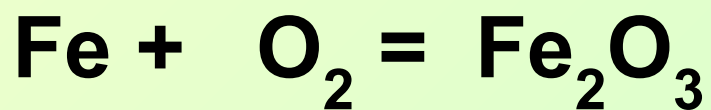


3

4







$\text{Fe}^0 - 3e = \text{Fe}^{+3}$  - ок-ие, в-ль

$\text{O}_2 + 4e = 2\text{O}^{-2}$  - в-ие, ок-

ль

4

3

# Расставьте коэффициенты методом электронного баланса

- $\text{Zn} + \text{HCl} = \text{ZnCl}_2 + \text{H}_2$
- $\text{Cu} + \text{HNO}_3 = \text{Cu}(\text{NO}_3)_2 + \text{NO}_2 + \text{H}_2\text{O}$ 
  - $\text{Al} + \text{O}_2 = \text{Al}_2\text{O}_3$
  - $\text{P} + \text{O}_2 = \text{P}_2\text{O}_5$

# Домашнее задание

- § 43
- Раб. тетр.: Стр. 162-163, упр. 12  
упр. 13(а-в)  
упр. 13